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have? 2. What is the
charge on a magnesium
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magnesium have to do to

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form such an ion, and why does it tend to do so? 3. How many valence electrons does a single neutral chlorine (Cl) atom have? 4.

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Name: _____ Date: _____ Hour: _____
Information: Energy
Levels and Sublevels ...

Considering your answer
to question 5, how many
electrons can fit in the
entire 5th energy level? can
(A_{Vel} is a ha AC5a)
5th ÷ 7. Write the
notation for an electron
spinning clockwise in a p
sublevel ...

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For each of the following
pairs of elements,

underline the one that
would replace the other
element in a compound.

- a. calcium, tin e. iron,
copper b. bromine,

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fluorine f. iodine,
chlorine c. aluminum,
potassium g. silver, lead
d. zinc, sodium 2.

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bond is polar covalent.

(b) $\Delta EN_{Cl-P} = EN_{Cl} - EN_P = 3.0 - 2.1 = 0.9$. Since $0.5 < \Delta EN_{Cl-P} < 1.7$, the PCl bond is polar covalent. (c)

ΔEN_{Cl-Li} Section 8.5
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and contrast polar . that

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shape of an s orbital?

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What is the relationship
between the size of an s
orbital and the principal
energy level in which it is
found? 3. What is the
shape of a p orbital? How

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many p orbitals are there
in a sublevel? 4.

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