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EQUATIONS. molarity =  $\frac{\text{mol solute}}{\text{L solution}}$ . 1 L =

1000 mL. The molarity of a solution is a ratio of the

moles of solute per liters of solution. The units for

molarity are written as mol/L or M. This measurement

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is used to perform stoichiometric calculations.

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= mol Water Concentrated solution is diluted with more solvent.  $M_1 \times V_1 = M_2 \times V_2$  = mol USEFUL EQUATIONS  $M_1 \times V_1 = M_2 \times V_2$  molarity =  $\frac{\text{molsolute}}{\text{L.solution}}$  1 L = 1000 ml. A solution can be made less concentrated in a process called dilution. This is accomplished by adding more solvent.

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USEFUL EQUATIONS  $\text{molarity} = \frac{\text{L. solution mol solute}}{\text{L. solution}}$

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1 L = 1000 mL The molarity of a solution is a ratio of the moles of.

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moles of solute per liters of solution. The units for molarity are written as mol/L or M. This measurement is

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As we learned previously, double replacement reactions involve the reaction between ionic compounds in solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because these reactions occur in aqueous solution, we can use the concept of molarity to directly calculate the number of moles of reactants or products that will ...

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new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics. The reorganization presents many core concepts in the first five chapters, preparing students for later chapters. In addition, the author has added vignettes throughout the chapters referring to health, technology, the environment and society as well as to specific tools of direct use to students.

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