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Edition

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Pogil Chemistry Answers

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Chemistry Answers

Calculating Ph Pogil Answer

Key $\text{pH} = 14.000 - \text{pOH} =$

$14.000 - 2.187 = 11.813$ 4)

-A solution is created by measuring 3.60×10^3 moles of NaOH and 5.95×10^{-4} moles of HCl into a container and then water is added until the final volume is 1.00 L. Calculating pH and pOH worksheet

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A Calculating pH - Manasquan
Public Schools pOH =
 $-\log[\text{OH}^-] = -\log(6.2 \times 10^{-5})$
 $= 4.21$ pH = 14.00 - pOH =
 $14.00 - 4.21 = 9.79$ 6) +A
-solution with a H
concentration of 1.00×10^{-7}
M is said to be neutral.
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of ph and poh is always 14
this is because the product
of proton concentration and
hydroxide concentration must
always equal the equilibrium
constant for the ionization
of water which is equal to
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Answers Pogil A Calculating pH - Manasquan Public Schools
 $\text{pOH} = -\log[\text{OH}^-]$
 $= -\log(6.2 \times 10^{-5}) = 4.21$
 $\text{pH} = 14.00 - \text{pOH} = 14.00 - 4.21 = 9.79$
6) +A - solution with a H concentration of 1.00×10^{-7} M is said to be neutral.

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